the molecular weight of the precipitate was calculated. The values so obtained, 190.8, 190.4 and 188.2, differ from the theoretical for cuprous xanthate ($CuSCSOC_2H_5$), 184.7, by only two or three per cent., the discrepancy being attributed to occlusion of dixanthogen. As the molecular weight of cupric xanthate ($Cu(SCSOC_2H_5)_2$), 305.9, is very much larger, the precipitate is undoubtedly the cuprous salt.

CONTRIBUTION FROM THE
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The Melting Point of Normal Butyl Ether.—Incidental to some other research the melting point of normal butyl ether has been determined, we believe, for the first time.

Eastman Kodak Company pure normal butyl ether was further purified by distillation in a small laboratory column, the middle third being used.

The apparatus used in this measurement consisted of a copper calorimeter wound with a heater and supplied with a copper–constantan thermocouple. This was hung by threads within a heavy copper cylinder, which was also provided with heater and thermocouple. The whole was suspended in a silvered glass tube connected to a mercury vapor pump to provide thermal insulation.

The apparatus was cooled below the melting point of the normal butyl ether and evacuated. The calorimeter was then warmed gradually by means of the heating coil. The ether melted very sharply at 177.8 ± 0.2°K., the sharpness of the melting point being sufficient evidence of the purity of the sample.

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